

CALORIMETRY

$$q_w = m C_w \Delta T$$

find q_w , change sign
for reaction

$$-q_w = \Delta H_{rxn}$$

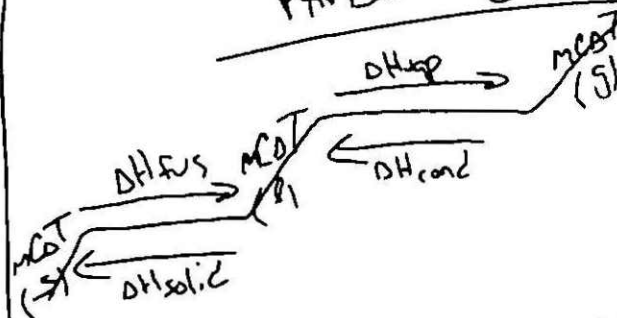
Calorimetry Probs Handout

HESS'S LAW

Rearrange & combine rxns
to find ΔH for unknown rxn

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(-750 kJ)

PHASE CHANGE



$$\Delta H_{PT} = mCoT_{(s)} + \Delta H_{fus} + mCoT_{(l)} + \Delta H_{vap} + mCoT_{(g)}$$

(s) or cond (g)

p. 4 packet

ΔH_f°

$$\Delta H^\circ = \sum \Delta H_{f,prod}^\circ - \sum \Delta H_{f,react}^\circ$$

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